

## ALTERNATIVE USE OF THE HYDROGEL 3005K4® COPOLYMER WITH LIGNINOLYTIC ENZYMES FOR TREATING CONTAMINATED SOIL

Uso alternativo del copolímero de hidrogel 3005K4® con enzimas ligninolíticas para el tratamiento de suelo contaminado

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Key words: superabsorbent polymer, bioremediation, hydrophilic polymers, *Trametes versicolor*.

### ABSTRACT

In this study, we found that a copolymer of acrylamide and potassium acrylate (Hydrogel 3005K4®) in two particle sizes (1-2 mm and 2-4 mm) could absorb conditioned Galhaup media recovered from *Trametes versicolor* growth containing ligninolytic enzymes. It was found that the most efficient absorption capacity of the copolymer in the presence of the conditioned medium was using a 0.2:10 weight/volume (w/v) copolymer-conditioned medium ratio for both particle sizes. This ratio is a quarter of its absorption capacity in distilled water. A different desorption behavior of the medium was also observed, depending on whether the copolymer was exposed to the soil, water, or air. The ligninolytic enzyme media absorbed by the Hydrogel 3005K4® showed slower desorption of conditioned medium in soil for 10 days and faster desorption in air and water between 24 and 72 h. The enzymes were capable to preserve the activity after absorption. The versatility of this vector holding the enzymes may be attractive for reduce the presence of organic compounds (pesticides, polycyclic aromatic hydrocarbons, synthetic and natural dyes) on soil as part of a bioremediation technique, with a possibility of being reused.

Palabras clave: polímero superabsorbente, biorremediación, polímeros hidrofílicos, *Trametes versicolor*.

### RESUMEN

En este estudio se encontró que un copolímero de acrilato de acrilamida y potasio (Hidrogel 3005K4®) en dos tamaños de partícula (1-2 mm y 2-4 mm), puede absorber medio condicionado de Galhaup recuperado del crecimiento de *Trametes versicolor* que contiene enzimas ligninolíticas. Se encontró que la capacidad de absorción más eficiente del copolímero en presencia del medio condicionado fue empleando una proporción de 0.2:10 peso/volumen (p/v) de copolímero-medio condicionado para ambos tamaños de partícula. Esta proporción es un cuarto de su capacidad de absorción en agua destilada. Se encontró un comportamiento diferente de desadsorción del medio enzimático contenido en el hidrogel, dependiendo de si el copolímero estaba o no

expuesto al suelo, agua o aire. El medio con las enzimas ligninolíticas absorbido por el Hidrogel 3005K4® mostró una desadsorción lenta del medio condicionado en el suelo a los 10 días y una rápida desadsorción en aire y agua entre las 24 y 72 h. Las enzimas fueron capaces de continuar con su actividad después de la absorción. La versatilidad de este vector que contiene las enzimas puede ser atractiva para reducir la presencia de compuestos orgánicos (plaguicidas, hidrocarburos aromáticos policíclicos, colorantes sintéticos y naturales) en el suelo como parte de las técnicas de biorremediación.

## INTRODUCTION

Hydrogels are highly hydrophilic polymers that are capable of absorbing and releasing substantial amounts of fluids because of their elastic three-dimensional networks and the presence of functional groups such as OH-, COOH-, and CONH- (Ahmed 2015). The degree of swelling, absorption capacity, and permeability are qualities of hydrogels, and these abilities depend on the type of material with which the hydrogel has been manufactured (Bai et al. 2010, Ekebafe 2011, Asgher et al. 2012, Ahmed 2015). Some factors can affect the degree of swelling in hydrogels, such as pH, ionic strength, temperature, and electromagnetic radiation (Chen and Zhao 2000, Ekebafe 2011, Ahmed 2015). Most applications of hydrogels are focused on the field of medicine and the controlled release of drugs. Currently, research has been developed for the application of hydrogels in agriculture, for the partial release of pesticides, in bioremediation by creating fungal inoculums, and in the capture of metal ions for automotive and industrial use waters (Galhaup et al. 2002, Escobar et al. 2003, Farrel et al. 2013, Ahmed 2015). Super-absorbent polymers are hydrogels that have been widely accepted for their liquid capture capacity and their wide application in various agricultural, pharmaceutical, and construction sectors, and can reach an absorption of up to 500 times their weight (Hublik and Schinner 2000, Kabiri and Zohuriaan-Mehr 2004, Kumar et al. 2009, Li 2010).

Molecular immobilization is a technique that allows the industrial application of enzymes because it improves stability, activity, and handling, which translates into prolonged activity times (Nannipieri and Bollag 1991, Ortiz et al. 2006, Liu et al. 2009, Muñoz and Zuluaga 2009). However, this process is carried out with polymerizations prepared at the time of the immobilization, and to date, there are no reports of using prefabricated copolymers for the immobilization of nutrient medium conditioned with enzymatic activities. Thus, this study aimed to develop an alternative to use Hydrogel 3005K4® as

an immobilizing vehicle for a conditioned medium containing ligninolytic enzymes with application potential in bioremediation. The latter can be achieved thanks to the high absorption of the copolymer in liquid media and the interactions between the positive charges, specifically of the enzyme laccase. This enzyme has the most tolerance to changes in pH and temperature, generating an interaction with the negative charges of the networks formed by the acrylamide acrylate potassium copolymer. Once the copolymer absorbs the conditioned medium, it is kept in the three-dimensional network. As the copolymer dehydrates or is replaced with water, the conditioned medium containing the enzymes of interest is released.

## MATERIALS AND METHODS

### Copolymer

Hydrogel 3005K4®, a copolymer of acrylamide and potassium was used with two particle sizes, 1-2 mm and 2-4 mm. Hydrogel 3005K4® has an absorption capacity of 270 to 400 times its weight in deionized water. In water with 1000 ppm of NaCl only, the absorption capacity is 150 to 270 times its weight, and in the presence of different electrolytes, the absorption capacity is 100 to 150 times its weight.

### Preparation of the conditioned medium

The ligninolytic fungus *Trametes versicolor* CDBB-H-1051 was obtained from the microbial collection of the Departamento de Biotecnología y Bioingeniería at Centro de Investigaciones y Estudios Avanzados of the Instituto Politécnico Nacional, Mexico. Petri dishes with potato dextrose agar (PDA) were inoculated with the fungi for propagation of the mycelium and were incubated at  $30 \pm 1$  °C for 10 days under natural light. The fungus was grown in liquid medium. The inoculum was prepared with 20 squares of  $4.2 \times 4.21$  mm ligninolytic fungi grown in PDA medium added to 500 mL Erlenmeyer flasks with yeast-peptone-glucose (YPG) medium. The flasks

were incubated with continuous stirring at 100 rpm at  $30 \pm 1$  °C for 10 days under natural light. Subsequently, the biomass was filtered and homogenized for 5.0 s with an electric hand blender in a sterile beaker. A 5.0% volume/volume (v/v) mixture of the homogenized fungus was added to the modified Galhau medium (Pouci 2008) and allowed to incubate for 30 days at  $30 \pm 1$  °C and 100 rpm under natural light. Next, the medium was filtered through Whatman® 1 filter paper using a vacuum to separate the fungus from the culture medium, and the recovered liquid was termed conditioned medium.

### Absorption of the conditioned Galhau medium

To test the absorption of the conditioned medium in the polymer, different ratios of acrylamide copolymer and potassium acrylate were tested with conditioned medium as shown in **table I**.

In addition, the time taken by the copolymer to absorb the conditioned medium and the amount of copolymer needed to absorb 100% of the conditioned medium was measured (Chen and Zhao 2000, Bai et al. 2010, Ekebafé 2011, Asgher et al. 2012, Ahmed 2015). Finally, the desorption percentage of conditioned medium at 35 °C was determined. Calculations of the percentage and degree of swelling and the percentage of absorption were performed using formulas 1, 2 and 3.

The percentage of hydration was determined from the following equation:

$$\% H = \frac{\text{Wet weight} - \text{Dry weight}}{\text{Dry weight}} * 100 \quad (1)$$

and the degree of swelling using the following formulae:

$$W = \frac{\text{Wet weight}}{\text{Dry weight}} \quad (2)$$

Percentage absorption:

$$\% A = \frac{(\text{Initial volume} - \text{final volume})}{\text{Initial volume}} * 100 \quad (3)$$

### Desorption of the conditioned Galhau medium

The desorption of the conditioned medium was determined when the copolymer was exposed to room temperature on a Whatman® 1 filter paper under vacuum conditions. The completely hydrated copolymer with the conditioned medium was placed on a filter paper 3.6 cm in diameter and was left at room temperature (25-30 °C) for 72 h. This procedure was performed in triplicate at 0, 3, 6, 24, 48, and 72 h to measure polymer moisture and weight. In addition, we determined whether there was desorption of the enzymes by performing a visual analysis (green color) of enzymatic activity on the filter paper using 1.8 mM 2,2'-azino-bis [3-ethylbenzothiazoline-6-sulfonic acid] (ABTS) directly. This analysis was based on the method described by Liu et al. in 2009, as there is no currently developed method to determine the enzymatic activity of this type of hydrogel.

To find the desorption of the conditioned medium from the copolymer in the soil, we used an 80 g soil sample from Centro Educativo de Riego of the División de Ciencias de la Vida (DICIVA) at the Universidad de Guanajuato. The sampling area was 330 m<sup>2</sup> and was located at a latitude of 20°44'39.2" N, longitude 10°11'939.4" W, and 1750 m asl with rainy and dry seasons throughout the year. The temperature varies from 7 °C to 32 °C. The soil was managed by conservation tillage for 12 years. The soil is classified as pelic vertisol with loamy-clayey texture according to the analyses performed at the Laboratorio de Tecnologías para la Sustentabilidad of the DICIVA. The sample was ground and sieved

**TABLE I.** VARIABLES AND LEVELS USED TO EVALUATE THE ABSORPTION OF THE CONDITIONED MEDIUM (CM) IN THE ACRYLAMIDE AND POTASSIUM ACRYLATE COPOLYMER.

Variables	Levels		
Copolymer + Conditioned medium (w/v)	0.05:10 (CM0)	0.1:10 (CM1)	0.2:10 (CM2)
Size of copolymer	1-2 mm 2-4 mm	1-2 mm 2-4 mm	1-2 mm 2-4 mm

w/v = weigh / volume

CM0 = copolymer-conditioned medium ratio 0.05: 10 w/v.

CM1 = copolymer-conditioned medium ratio 0.1: 10 w/v.

CM2 = copolymer-conditioned medium ratio 0.2: 10 w/v.

in # 20 mesh. One hundred grams of soil and 15 g of hydrated copolymer with conditioned medium were placed into a glass jar. The treatments (polymer sizes 1-2 mm and 2-4 mm) were maintained at  $28 \pm 2.0$  °C for 24 days. Samples were evaluated in triplicate for 0, 2.5, 10 and 24 days, and for each time, the soil moisture was evaluated to determine how it changed with the release of the conditioned medium.

Desorption of the conditioned medium in water was performed by placing the previously hydrated copolymer with the conditioned medium, with two different weights (1.5 and 3 g) in 10 mL of water. The weight of the copolymer was assessed on days 0, 2, and 5, while the absorbance of the aqueous medium surrounding the copolymer was quantified at 230 nm using a UV-Vis spectrophotometer on days 0 and 5. These measurements were conducted to elucidate the desorption kinetics of the enzymes.

ChemSketch program version 14.000 (Copyright © 1994-2013 Advanced Chemistry Development, Inc.) was used to identify possible interactions of the conditioned medium and the copolymer. Also, an analysis of variance (ANOVA) type III sums of squares was performed.

## RESULTS AND DISCUSSION

The absorption percentage obtained at various times of the conditioned medium showed a statistically significant correlation between the ratio of the copolymer-conditioned medium with the exposure time and copolymer particle size ( $p = 0.000$ ,  $p = 0.000$  and  $p = 0.026$  with  $\alpha = 0.05$ ). The controls showed more efficient absorption in less time and had a lower amount of 100% copolymer absorption at 6 h at a ratio of 0.1:10 w/v copolymer-water (**Table II**). For both copolymer particle sizes, improved absorption was detected when 0.2 g of copolymer per 10 mL of conditioned medium was added, resulting in an absorption ratio 50 times the weight of the copolymer. This ratio is less efficient than that of the control at 0.1 g of polymer with 10 mL of water. According to the supplier, this copolymer can absorb 270 to 400 times its weight in deionized water and 150 to 270 times its weight in water with 1000 ppm of NaCl. Therefore, its absorption capacity was much lower than that reported by the supplier, although the fluid conditions were not the same. This result was attributed to the presence of ions with positive

**TABLE II.** PERCENTAGE OF ABSORPTION DETECTED IN THE ACRYLAMIDE AND POTASSIUM ACRYLATE COPOLYMER IN THE PRESENCE OF THE CONDITIONED MEDIUM (CM) RECOVERED FROM *Trametes versicolor* USING TWO PARTICLE SIZES.

Polymer size	Time (h)	% Absorption		
		CM0	CM1	CM2
1-2 mm	0	4.33 ± 0.06	6.00 ± 0.00	8.33 ± 0.06
	3	31.33 ± 0.12	61.33 ± 0.12	100 ± 0.00
	6	32.67 ± 0.12	61.67 ± 0.07	100 ± 0.00
	24	32.67 ± 0.12	61.67 ± 0.07	100 ± 0.00
Cont* 1-2 mm	0	5.67 ± 0.00	12.67 ± 0.14	30.00 ± 1.73
	6	94.33 ± 0.07	100.00 ± 0.00	100.00 ± 0.00
	24	89.33 ± 0.00	100.00 ± 0.00	100.00 ± 0.00
2-4 mm	0	4.00 ± 0.00	3.67 ± 0.06	5.00 ± 0.14
	3	21.00 ± 0.00	46.33 ± 0.32	77.00 ± 0.14
	6	30.00 ± 0.20	54.67 ± 0.00	86.00 ± 0.00
	24	36.67 ± 0.14	60.67 ± 0.00	97.00 ± 0.00
Cont* 2-4 mm	0	6.67 ± 0.00	6.00 ± 0.14	5.50 ± 0.1
	6	86.67 ± 0.57	100.00 ± 0.00	100.00 ± 0.00
	24	92.00 ± 0.00	100.00 ± 0.00	100.00 ± 0.00

\*Cont = Water.

CM0 = copolymer-conditioned medium ratio 0.05: 10 w/v.

CM1 = copolymer-conditioned medium ratio 0.1: 10 w/v.

CM2 = copolymer-conditioned medium ratio 0.2: 10 w/v.

charges. Even with this absorption capacity, the costs of using this copolymer (approximately US \$ 7) are below the use of other materials such as alginates (approximately US \$ 170).

There are reports of absorption of various compounds, such as methylene blue, crystal violet and heavy metals or arsenic, by using specific designs of copolymers for each type of compound and manipulating the pH and temperature for their conditioning. Our experiment was conducted under pH 6.3 - 6.7 and room temperature (25 - 31 °C) (Zhen and Zhang 2008, Singh et al. 2010, Villarroel et al. 2010, Snoeck et al. 2014). To optimize the absorption processes, the conditions and compound to be absorbed must be evaluated, and a specific hydrogel or copolymer design must be assessed, which can result in considerable increases in costs. We used a single type of copolymer in different treatments to reduce costs.

In the analysis of hydration and degree of swelling, we found 150 to 300 times increased hydration and swelling in the CM0 and CM1 control treatments with respect to the treatments with the conditioned medium (**Table III**).

**Table III** shows that the degree of swelling and hydration of the copolymer were similar between the control and the samples with conditioned medium in the CM2 treatment (1 - 2 mm copolymer with conditioned medium  $4736.40 \pm 95.77$ ; control with water  $4678.44 \pm 106.61$ ; 2-4 mm copolymer with conditioned medium  $4331.40 \pm 57.15$ ; control with water  $4484.15 \pm 130.86$ ). This result shows that this relationship is ideal to recover the conditioned medium.

In the desorption process of the medium conditioned from the hydrogel, we found that both time and the copolymer's size had a statistically significant effect on the release of the conditioned medium ( $p = 0.0000$ ,  $p = 0.0013$ , respectively, with  $\alpha = 0.05$ ). In the case of desorption in soil, only time had a statistically significant effect, with day 10 being the maximum release point ( $p = 0.0000$  with  $\alpha = 0.05$ ). The desorption time of the conditioned medium decreased when exposed to the soil. In the case of soil, the desorption time was slower and began on day 10. The process was constant without desorption at three days of exposure for both particle sizes. Desorption reached a maximum desorption point of < 83%, as shown in **table IV**.

The release of the conditioned medium contained in the hydrogel in water was determined by changing the absorbance of the water. Due to the release of the characteristic color of the conditioned medium, it was

possible to determine the desorption of this medium in water. The absorbances correlated with the amount of copolymer, showing that the absorbance increased with the amount of copolymer (**Table V**).

In contrast, we detected that the copolymer, when in water, tended to release the conditioned medium and absorb the available water simultaneously. By this, the copolymer increased its weight by more than 70% when the ratio was 3 g of copolymer (with conditioned medium) to 10 mL of water and increased more than 85% when the ratio was 1.5 g of copolymer (with conditioned medium) to 10 mL of water (**Fig. 1**). From day 2, the weight of the copolymer varied between 0.02-0.4 g, showing that the release of the conditioned medium and the absorption of the water had ceased (**Fig. 1**).

One possible factor that affected the percentage of absorption of the conditionate medium is the presence of cations such as  $\text{Co}^+$  and  $\text{Mg}^+$ . The positive charges may compete for binding with the negative partial charges from the oxygen in the copolymer, causing interference between dipole-dipole bonds (**Fig. 2**). Additionally, the presence of enzymes that have a positive charge may compete with water and prevent the formation of hydrogen bonds with the molecules of the polymeric network that make up the copolymer, and can decrease their swelling and absorption capacities in a range of pH 5-9. Villarroel et al. (2010) used potentially biodegradable hydrogels obtained from polymerized and cross-linked acrylamide in the presence of the poly [3-hydroxybutyrate] biopolymer for the absorption of copper and nickel. They showed a high affinity and efficiency of absorption for copper molecules. This affinity results in a decrease in the capacity of the absorption process. During the desorption process on filter paper, enzymatic activity of laccase was detected, resulting in the characteristic green coloration of the ABTS oxidation reaction (**Fig. 3**).

## CONCLUSION

It has been proven that the use of copolymers for the immobilization of nutrient medium conditioned with viable enzymes is possible. The copolymer had reduced absorption capacity in the presence of the conditioned medium recovered from *T. versicolor* due to possible interactions between the ion-dipole or dipole-dipole forces of the copolymer and the  $\text{Mg}^+$ ,  $\text{Cu}^+$ , and enzymes with positive charges present in the conditioned medium. We found an absorption capacity of 50 times the weight of the copolymer.



**TABLE III.** HYDRATION CAPACITY AND DEGREE OF SWELLING OF THE ACRYLAMIDE AND POTASSIUM ACRYLATE COPOLYMER IN THE PRESENCE OF THE CONDITIONED MEDIUM RECOVERED FROM *Trametes versicolor* USING TWO PARTICLE SIZES.

Polymer size	Time (h)	% Hydration			Degree of swelling		
		CM0	CM1	CM2	CM0	CM1	CM2
1-2 mm	0	634.41 ± 49.82	456.17 ± 68.55	373.71 ± 23.34	7.34 ± 0.5	5.56 ± 0.69	4.74 ± 0.23
	3	5684.43 ± 99.69	5902.81 ± 147.91	4704.09 ± 71.60	57.84 ± 1.00	60.03 ± 1.48	48.04 ± 0.72
	6	5761.36 ± 59.00	5827.43 ± 75.72	4762.08 ± 103.66	58.61 ± 0.59	59.27 ± 0.76	48.62 ± 1.04
	24	5780.33 ± 87.90	5849.60 ± 193.62	4736.40 ± 95.77	58.80 ± 0.88	59.50 ± 1.94	48.36 ± 0.96
Cont* 1-2 mm	0	1008.42 ± 109.30	1059.84 ± 194.51	1234.44 ± 502.67	11.08 ± 1.09	11.60 ± 1.95	13.34 ± 5.03
	6	15069.67 ± 728.03	7658.67 ± 139.89	3770.02 ± 47.29	151.70 ± 7.28	77.59 ± 1.40	38.70 ± 0.47
	24	17046.06 ± 467.15	9392.13 ± 266.96	4678.44 ± 106.61	171.46 ± 4.67	94.92 ± 2.67	47.78 ± 1.07
2-4 mm	0	539.26 ± 71.62	239.93 ± 17.65	202.82 ± 56.77	6.39 ± 0.72	3.40 ± 0.18	3.03 ± 0.57
	3	3501.71 ± 13.30	4218.09 ± 315.49	3431.06 ± 74.52	36.02 ± 0.13	43.18 ± 3.15	35.31 ± 0.75
	6	4719.37 ± 364.23	4867.23 ± 305.84	4124.11 ± 79.52	44.86 ± 6.32	49.67 ± 3.06	42.24 ± 0.80
	24	5622.87 ± 432.35	5448.30 ± 181.25	4331.40 ± 57.15	57.23 ± 4.32	55.48 ± 1.81	44.31 ± 0.57
Cont* 2-4 mm	0	707.37 ± 123.82	398.19 ± 154.94	161.90 ± 26.62	8.07 ± 1.24	4.98 ± 1.55	2.62 ± 0.27
	6	12962.33 ± 963.47	7785.47 ± 418.44	3845.54 ± 121.25	130.62 ± 9.63	78.85 ± 1.98	39.46 ± 1.21
	24	14898.65 ± 207.61	9354.97 ± 197.60	4484.15 ± 130.86	157.64 ± 13.33	94.55 ± 1.98	45.84 ± 1.31

\*Cont = Water

CM0 = copolymer-conditioned medium ratio 0.05: 10 w/v

CM1 = copolymer-conditioned medium ratio 0.1: 10 w/v

CM2 = copolymer-conditioned medium ratio 0.2: 10 w/v

**TABLE IV.** PERCENTAGE RELEASE OF CONDITIONED MEDIUM WHEN THE ACRYLAMIDE AND POTASSIUM ACRYLATE COPOLYMER WAS EXPOSED TO AIR AND SOIL.

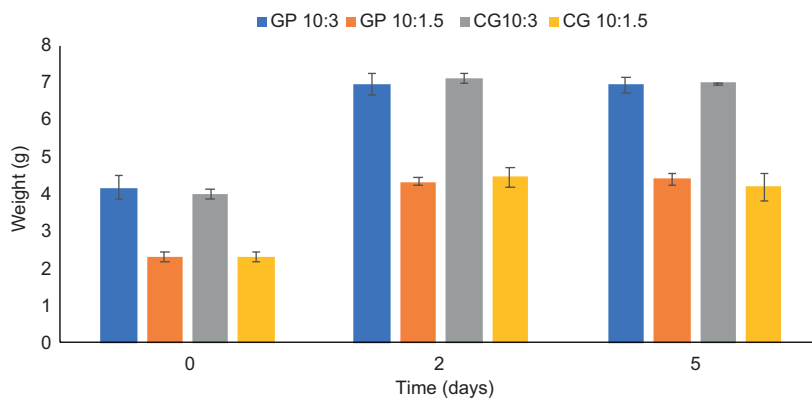
Copolymer	Time (h)	% of conditioned medium released	
		CS	CL
Exposed to Air	0	0.00 ± 0.00	0.00 ± 0.00
	3	16.33 ± 0.60	10.85 ± 0.05
	24	51.72 ± 5.16	42.51 ± 4.41
	48	95.78 ± 0.90	72.03 ± 1.82
	72	97.1 ± 0.13	96.87 ± 0.22
Exposed to Soil	0	0.00 ± 0.00	0.00 ± 0.00
	48	41.90 ± 2.53	41.16 ± 3.76
	120	52.61 ± 1.36	51.16 ± 1.71
	240	74.74 ± 4.13	79.70 ± 0.72
	576	79.05 ± 2.87	82.56 ± 6.32

CS = 1-2 mm copolymer (small copolymer).  
CL = 2-4 mm copolymer (large copolymer).

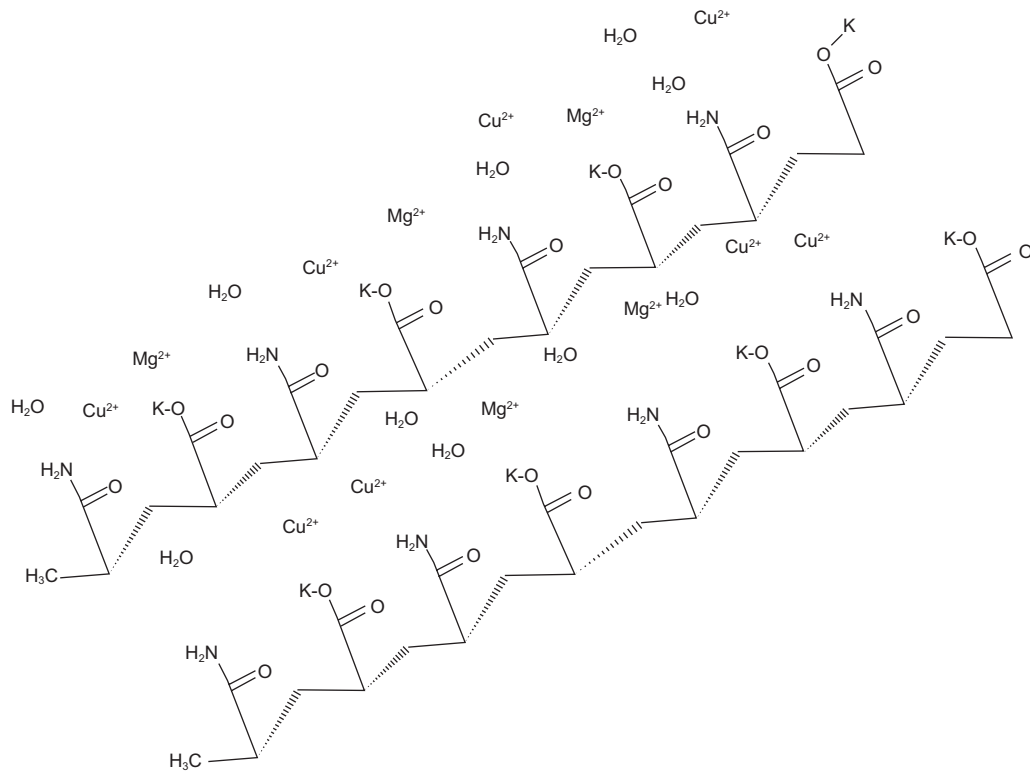
**TABLE V.** RELEASE OF LACCASE WHEN THE ACRYLAMIDE AND POTASSIUM ACRYLATE COPOLYMER WAS EXPOSED TO WATER BY MEASURING THE ABSORBANCE AT 230 NM.

Copolymer	Time	Absorbance	
		CS	CL
Water 10:3	0	0.19 ± 0.02	0.15 ± 0.03
	120	0.75 ± 0.01	0.93 ± 0.02
Water 10:1.5	0	0.10 ± 0.01	0.10 ± 0.01
	120	0.37 ± 0.00	0.45 ± 0.06

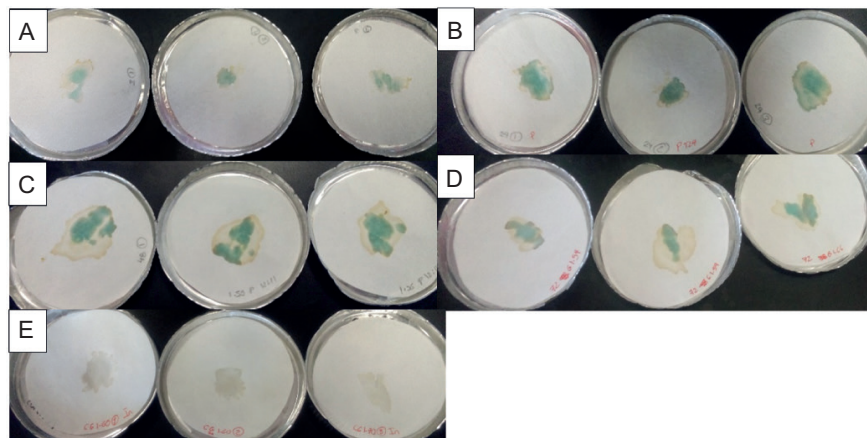
10:3 = ratio of 10 mL of water to 3 g of copolymer with conditioned medium.  
10:1.5 = ratio of 10 mL of water to 1.5 g of copolymer with conditioned medium.  
CS = Small copolymer  
CL = Large copolymer



**Fig. 1.** Behavior of the copolymer with conditioned medium when immersed in water (CP = copolymer 1-2 mm; CG = copolymer 2-4 mm; 10:3 = ratio of 10 mL of water to 3 g of copolymer with conditioned medium; 10:1.5 = ratio of 10 mL of water to 1.5 g of copolymer with conditioned medium).



**Fig. 2.** Proposed behavior of the interaction of the conditioned medium with the copolymer molecule.



**Fig. 3.** Enzymatic activity of laccase released from the acrylamide and potassium acrylate copolymer on filter paper after 72 h when reacting with ABTS (1.8 mM), A) 3 h, B) 24 h, C) 48 h, D) 72 h, E) control with water only. ABTS = 2,2'-azino-bis[3-ethylbenzothiazoline-6-sulfonic acid].



In the desorption process, the conditioned medium was released into water and soil and exposed to the environment, and we found that the copolymer had an affinity for water under the condition tested. Laccase activity was seen visually during desorption on filter paper. However, the initial and final enzyme units still need to be determined due to the conditions of the assay. These results help to consider whether this mixture could help to release the enzymes during a bioremediation process and at the same time, to keep the humidity on the plant.

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